

- 1) From each one of your graphs that you constructed in excel, calculate the number of moles of the gas in your experiment. Show your work.

HINTS: For Charles law and Gay-Lussacs law, this will involve relying on our skills from ph201 (i.e. extracting information from a linear plot). What does the slope of each graph represent in terms of the ideal gas law ($P V = n R T$).

For Boyles law, you used a power fit, which resulted in an equation that looks like the following...
 $y = A x^c$. Hopefully your c value is very close to -1. With $c=-1$, can you now relate what the value of A is in terms of the ideal gas law?